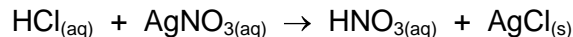
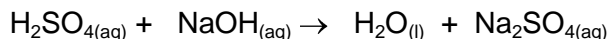


Stoichiometry Practice with Molarity Name: \_\_\_\_\_ Per: \_\_\_ Date: \_\_\_\_\_

1. You measured out a 25.0 mL sample of HCl, then added silver nitrate solution and filtered off the white precipitate of silver chloride, AgCl. If the dry precipitate weighed 0.3537 g, what was the molarity of the hydrochloric acid?

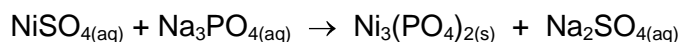


2. Consider the reaction of sulfuric acid, H<sub>2</sub>SO<sub>4</sub>, with sodium hydroxide, NaOH. Note the typical reaction of an acid with a base producing water and a salt.



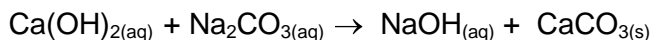
Suppose a beaker contains 35.0 mL of 0.175 M H<sub>2</sub>SO<sub>4</sub>. How many milliliters of 0.250 M NaOH must be added to react completely with the sulfuric acid?

3. Nickel sulfate, NiSO<sub>4</sub>, reacts with trisodium phosphate, Na<sub>3</sub>PO<sub>4</sub>, to give a pale yellow-green precipitate of nickel phosphate, Ni<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub> and a solution of sodium sulfate, Na<sub>2</sub>SO<sub>4</sub>.



How many milliliters of 0.375 M NiSO<sub>4</sub> will react with 45.7 mL of 0.265 M Na<sub>3</sub>PO<sub>4</sub>?

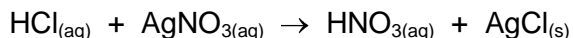
4. Sodium hydroxide, NaOH, may be prepared by reacting calcium hydroxide, Ca(OH)<sub>2</sub>, with sodium carbonate, Na<sub>2</sub>CO<sub>3</sub>.



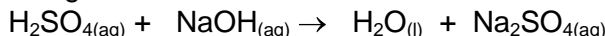
How many milliliters of 0.0150 M Ca(OH)<sub>2</sub> will react with 2.55 g Na<sub>2</sub>CO<sub>3</sub>?

Stoichiometry Practice Name: \_\_\_\_\_ Per: \_\_\_ Date: \_\_\_\_\_

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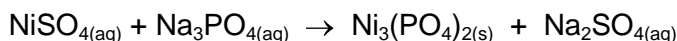


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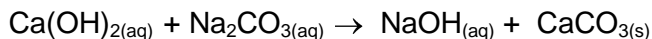
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